| الاسم: | مسابقة في مادة الفيزياء |
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| الرقم: | المدة: ساعة واحدة |

## The exam is formed of three exercises. The use of non-programmable calculators is recommended.

## I-(08 points)

## Mechanical Energy

From a point $O$, found at 25 m above the ground, Tarek throws a stone $(S)$, of mass 0.1 kg , vertically upwards with a speed of $20 \mathrm{~m} / \mathrm{s}$.
$(S)$ thus reaches a highest point $A$ then falls down to the ground at $B$.
Neglect air resistance and take $g=10 \mathrm{~m} / \mathrm{s}^{2}$.
The horizontal level of the ground $B$ is taken as the gravitational potential energy reference for the system [ $(S)$, Earth].

1. Calculate, at the instant of shooting at point $O$ :
a) the kinetic energy of $(S)$.
b) the gravitational potential energy of the system [( $S$ ), Earth].
c) the mechanical energy of the system [ $(S)$, Earth].
2. a) The mechanical energy of the system [ $(S)$, Earth] is conserved. Why?
b) What is the speed of $(S)$ at the highest point $A$ ? Deduce the kinetic energy of $(S)$ at $A$.

c) Determine the gravitational potential energy of the system [ $(S)$,

Earth] at $A$ and deduce the height of $A$ above the ground.
3. $(S)$ meets the ground at $B$. Determine, using the conservation of mechanical energy, the kinetic energy and the speed of $(S)$ at $B$.

## II-(06 points) <br> Nuclear Fusion

In young stars, as our Sun, two hydrogen nuclei $\left({ }_{1}^{1} H\right)$ fuse to form deuterium nucleus $\left({ }_{1}^{2} H\right)$. Two deuterium nuclei fuse to give a tritium $\left({ }_{1}^{3} H\right)$. However, from the energetic point of view, the most interesting fusion reaction is that one corresponding to a deuterium nucleus $\left({ }_{1}^{2} H\right)$ and a tritium nucleus $\left({ }_{1}^{3} \mathrm{H}\right)$. This reaction produces a helium nucleus $\left({ }_{2}^{4} \mathrm{He}\right)$ and a neutron.

## Given the masses of the nuclei in the atomic mass unit:

$$
\begin{aligned}
& m\left({ }_{1}^{1} H\right)=1.0073 u ; \\
& m\left({ }_{1}^{3} H\right)=3.0155 u ; \\
& 1 u=1.66 \times 10^{-27} \mathrm{~kg} ;
\end{aligned}
$$

$$
m\left({ }_{1}^{2} H\right)=2.0136 u ;
$$

$$
m\left({ }_{2}^{4} \mathrm{He}\right)=4.0015 u ;
$$

$$
m\left({ }_{0}^{1} n\right)=1.0087 u ;
$$

$$
c=3 \times 10^{8} \mathrm{~m} / \mathrm{s} .
$$

## Questions:

1. $\left({ }_{1}^{1} H\right),\left({ }_{1}^{2} H\right)$, and $\left({ }_{1}^{3} H\right)$ are isotopes of the same element. Why?
2. What is the number of neutrons in the tritium nucleus?
3. The first fusion reaction described in the text is of form: ${ }_{1}^{1} H+{ }_{1}^{1} H \longrightarrow{ }_{1}^{2} H+{ }_{Z}^{A} X$

Determine $A$ and $Z$ and give the name of $X$.
4. One of the reactions mentioned in the text produces a helium nucleus $\left({ }_{2}^{4} \mathrm{He}\right)$.
a) Write the corresponding equation of the reactions.
b) Calculate, in $u$ then in kg , the mass defect in this reaction.
c) Calculate, in joules, the energy liberated by this reaction.

## III-(06 points)

## Solar System

Certain characteristics of the planets of our solar system are summarized in the following table:

| Planet | Average distance from the Sun (AU) | Speed of the planet along its orbit (km/s) | One revolution around the Sun (year) | Density ( $\mathrm{g} / \mathrm{cm}^{3}$ ) | Chemical composition of the atmosphere | Surface temperature $\left({ }^{\circ} \mathrm{C}\right)$ |
| :---: | :---: | :---: | :---: | :---: | :---: | :---: |
| Mercury | 0.38 | 47.9 | ? | 5.43 | $\mathrm{H}_{2}$, He | -170 to 400 |
| Venus | 0.72 | 35 | 0.61 | 5.24 | $\mathrm{CO}_{2}(95 \%)$ | 480 |
| Earth | 1 | 29.8 | 1 | 5.51 | $\mathrm{N}_{2}, \mathrm{O}_{2}$ | 22 |
| Mars | 1.52 | 24.1 | 1.88 | 3.94 | $\mathrm{CO}_{2}$ | -170 to 35 |
| Jupiter | 5.20 | 13.1 | 11.86 | 1.33 | $\mathrm{H}_{2}$, He | -150 |
| Saturn | 9.53 | 9.6 | 29.45 | 0.69 | $\mathrm{H}_{2}$, $\mathrm{He}, \mathrm{CH}_{4}$ | -180 |
| Uranus | 19.19 | 6.8 | 84 | 1.30 | $\mathrm{H}_{2}, \mathrm{He}, \mathrm{CH}_{4}$ | -500 |
| Neptune | 30 | 5.4 | 164 | 1.76 | $\mathrm{H}_{2}$, $\mathrm{He}, \mathrm{CH}_{4}$ | -210 |
| Pluto | 39.53 | 4.8 | 247.7 | 2 | $\mathrm{N}_{2}$ | -230 |

## Questions:

1. Scientists classify the planets into two groups: the internal group and the external group.

The planets: Mercury, Venus, Earth, and Mars form a group.
a) Which group?
b) Point out, from the table, a physical quantity that characterizes their constitution.
2. The time taken by a planet to perform one complete revolution around the Sun is expressed in "years".
What does "year" stand for? Justify your answer using the table.
3. The distance Earth-Sun, of value $150 \times 10^{6} \mathrm{~km}$, is called the astronomical unit (AU).

Calculate, in km, the distance Uranus-Sun.
4. Compare the temperature at the surface of Mercury and of Venus and their average distances from the Sun, answer the following questions:
a) What is the strange observation noted about Venus?
b) Due to what is this observation?
5. What conclusion may be drawn as a result of comparing:
a) The period of revolution of the planets with their average distance from the Sun?
b) The speeds of the planets with their average distances from the Sun?

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| الالاسم: | مسابقة في مادة الفيزياء المدة ساعة | مشروع معيار التصحيح |

First Exercise (8 points)

| Part | Answer | Mark |
| :---: | :---: | :---: |
| 1.a | K. $\mathrm{E}_{\mathrm{O}}=\frac{1}{2} \mathrm{mV} V_{\mathrm{O}}^{2}=\frac{1}{2}(0.1)(20)^{2}=20 \mathrm{~J}$ | 1 |
| 1.b | G. P. $\mathrm{E}_{\mathrm{O}}=\mathrm{mgh}=(0.1)(10)(25)=25 \mathrm{~J}$ | 1 |
| 1.c | M. $E_{O}=$ K. $E_{O}+$ G.P. $E_{O}=20+25=45 \mathrm{~J}$ | 1 |
| 2.a | M.E is conserved since there is no air resistance. | 0.25 |
| 2.b | $\mathrm{V}_{\mathrm{A}}=0 \mathrm{~m} / \mathrm{s}$ then $\mathrm{K} . \mathrm{E}_{\mathrm{A}}=\frac{1}{2} \mathrm{mV} \mathrm{V}_{\mathrm{A}}^{2}=0 \mathrm{~J}$ | 2 |
| $2 . \mathrm{c}$ | By applying the law of conservation of mechanical energy: <br> M. $\mathrm{E}_{\mathrm{o}}=\mathrm{M} . \mathrm{E}_{\mathrm{A}} \Rightarrow 45=$ K. $\mathrm{E}_{\mathrm{A}}+$ G.P. $\mathrm{E}_{\mathrm{A}}$ (K. $\mathrm{E}_{\mathrm{A}}=0 \mathrm{~J}$ maximum position) <br> G. P. $E_{A}=45 \mathrm{~J} \Rightarrow$ G. P. $\mathrm{E}_{\mathrm{A}}=\mathrm{mgh}_{\mathrm{A}} \Rightarrow \mathrm{h}_{\mathrm{A}}=\frac{\text { G.P.E }}{\mathrm{mg}}=\frac{45}{0.1 \times 10}=45 \mathrm{~m}$ | 2 |
| 3 | By applying the law of conservation of mechanical energy: <br> M. $\mathrm{E}_{\mathrm{o}}=\mathrm{M} . \mathrm{E}_{\mathrm{B}} \Rightarrow 45 \mathrm{~J}=\mathrm{K} \cdot \mathrm{E}_{\mathrm{B}}+\mathrm{G} . \mathrm{P} \cdot \mathrm{E}_{\mathrm{B}}$ (G.P. $\mathrm{E}_{\mathrm{B}}=$ <br> 0 J since $B$ is on the reference) $K . E_{B}=45 \mathrm{~J} \Rightarrow \frac{1}{2} m v_{B}^{2}=45 \Rightarrow \frac{1}{2}(0.1) v_{B}^{2}=45 \Rightarrow v_{B}=30 \mathrm{~m} / \mathrm{s}$ | 2.25 |

## Second Exercise (8 points)

| Part | Answer | Mark |
| :---: | :---: | :---: |
| 1 | They are isotopes of the element hydrogen since they are nuclides that have the same charge number $\mathrm{Z}=1$ and different mass number A . | 0.5 |
| 2 | $\mathrm{N}=\mathrm{A}-\mathrm{Z}=3-1=2$ | 0.50 |
| 3 | By applying the law of conservation of the mass number: $1+1=2+A \Rightarrow$ $A=0$ <br> By applying the law of conservation of the charge number: $1+1=1+\mathrm{Z} \Rightarrow$ $\mathrm{Z}=1$ <br> Then the particle X is positron ${ }_{+1}^{0} \mathrm{e}$ | 1.5 |
| 4.9 | The equation can be written as: ${ }_{1}^{2} \mathrm{H}+{ }_{1}^{3} \mathrm{H} \rightarrow{ }_{2}^{4} \mathrm{He}+{ }_{\mathrm{Z}}^{\mathrm{A}} \mathrm{X}$ <br> By applying the law of conservation of the mass number: $2+3=4+A \Rightarrow$ $A=1$ <br> By applying the law of conservation of the charge number: $1+1=2+Z \Rightarrow$ $Z=0$ <br> Then the particle X is a neutron ${ }_{0}^{1} n$ <br> Fusion reaction ${ }_{1}^{2} \mathrm{H}+{ }_{1}^{3} \mathrm{H} \rightarrow{ }_{2}^{4} \mathrm{He}+{ }_{0}^{1} \mathrm{n}$ | 0.5 |
| 4.b | $\begin{aligned} & \Delta \mathrm{m}=\mathrm{m}_{\mathrm{b}}-\mathrm{m}_{\mathrm{a}}=\left[\mathrm{m}\left({ }_{1}^{2} \mathrm{H}\right)+\mathrm{m}\left({ }_{1}^{3} \mathrm{H}\right)\right]-\left[\mathrm{m}\left({ }_{2}^{4} \mathrm{He}\right)+\mathrm{m}\left({ }_{0}^{1} \mathrm{n}\right)\right] \\ & =[2.0136 \mathrm{u}+3.0155 \mathrm{u}]-[4.0015 \mathrm{u}-1.0087 \mathrm{u}] \\ & =0.0189 \mathrm{u}=0.0189 \times 1.66 \times 10^{-27}=3.137 \times 10^{-29} \mathrm{~kg} \end{aligned}$ | 1.75 |
| 4.c | $\mathrm{E}=\Delta \mathrm{mc}^{2}=\left(3.137 \times 10^{-29}\right)\left(3 \times 10^{8}\right)^{2}=28.237 \times 10^{-13} \mathrm{~J}$ | 1.25 |

Third Exercise ( 6 points)

| Part | Answer | Mark |
| :---: | :--- | :---: |
| 1.a | Inner planets or terrestrial planets | 0.5 |
| $\mathbf{1 . b}$ | From the table, the planets Mercury, Venus, Earth, and Mars have respectively <br> as densities $5.43 ; 5.24 ; 5.57$ and $3.94 \mathrm{~g} / \mathrm{cm}^{3}$. The density being almost the same <br> for the 4 planets, we can then consider it as characteristic physical quantity of the <br> inner planets. | 0.5 |
| $\mathbf{2}$ | In the table, the duration of revolution of Earth around the Sun is equal to one <br> unit. <br> Thus, the year that is taken as unit is the Earth year. | 1.25 |
| $\mathbf{3}$ | Distance Uranus-Sun=19.9AU=19.9×150×10 $=2.985 \times 10^{9} \mathrm{~km}$ | 1 |
| 4.a | From the table, we notice that the temperature of the surface of Venus is higher <br> than that of the surface of Mercury although Venus is farther from the Sun than <br> Mercury. | 0.75 |
| 4.b | Atmosphere of Venus made up of carbon dioxide which causes Greenhouse <br> effect whereas Mercury's atmosphere is thin and greenhouse effect doesn't take <br> place. | 0.5 |
| 5.a | From the table, the period of a planet around the Sun increases when the <br> distance of this planet from the Sun increases. | 0.75 |
| 5.b | From the table, the speed of a planet around the Sun increases when the <br> distance of this planet from the Sun decreases. | 0.75 |

